MCQ [illegible]

(A) Theory
(B) Hydrogen bond
(C) Benzene
(D) Coke
(E) Ether
(F) Acetylene
(G) Glycerol
(H) Gasoline
(I) Alcohol
(J) Ethanol
(K) Methane
(L) Sucrose
(M) Water
(N) Acetic
(O) Organized Level
(P) Molecular Shell
(Q) Regulation Level
(R) High Specific Heat
(S) Biotic
(T) Atomic Mass
(U) Mole Ratio
(V) Atomic Number
(W) Isotope
(X) Air
(Y) Hydrogen Ion
(Z) Cohesion
(AA) Evaporation
(AB) Adhesion
(AC) Atomic Number
(AD) Isotopes
(AE) Energy
(AF) Cellular Level
(AG) Atomic Organization
(AH) Covalent Bond
(AI) Organelles
(AJ) Ionic Bond
(AK) Radioactive
(AL) pH
(AM) Covalent Bond
(AN) Colloid
(AO) Hydrogen
(AP) Hydrogen Molecule
(AQ) Ion
(AR) Homeostasis
(AS) Isotopic Mass
(AT) Energy Level (shell)
(AU) Hydrogen
(AV) Sugar

1. _____ Level of hierarchy which includes tissues, organs, organ systems, and organisms.

2. _____ The frequency of a wave that is when frequency is twice that of a pendulum. Do not use.

Matchstick - Chapter 1 and 2

3. _____ A process could be not a hypothesis.

4. _____ The number of protons and neutrons.

5. _____ The component which makes a substance acidic.

6. _____ Organization of the biological world which shows how each level builds on the level below.

7. _____ A substance that contains more OH- than H+ in solution with a high pH.

8. _____ The atomic property with an charge.

9. _____ A region in which an electron is most likely to be found.

10. _____ The way in which life is unique.

11. _____ The scientist who proposed evolution.

12. _____ A molecule explanation.

13. _____ A structure that forms around an ion when a compound is water. Forces around ions form hydrogen bonds.

14. _____ A bond which forms between atoms with a partial charge due to the difference in electronegativity.

15. _____ A form of energy commonly used as fuel.

16. _____ The affinity for electrons by an atom nucleus.

17. _____ Level of hierarchy which includes populations, species, ecosystems, communities, and biosphere.

18. _____ The component that makes a substance more basic.

19. _____ Bonds into changing form different from Ionic bond pairs.

20. _____ A specification part of the cell.

21. _____ The attraction of one water molecule to another water molecule.

22. _____ The amount of energy an electron has to relate to the nucleus of an atom.

23. _____ The appearance of a new species due to a split in a population caused by reproductive isolation.

24. _____ Required to organisms to grow, reproduce, avoid predators, fight disease, and maintain homeostasis.

25. _____ Characteristics that are not unique to water.

26. _____ A type of covalent bond that involves the equal sharing of electrons.

27. _____ A substance particle with a positive charge.

28. _____ A metal explanation.

29. _____ The measure of the concentration.

30. _____ Bond which is the result of sharing an accepting electron.

31. _____ Indicates the number of protons.

32. _____ An acidic substance which is commonly used in blood buffering.

33. _____ A bond which forms in the result of sharing electrons.

34. _____ Maintaining a relatively balanced internal environment.

35. _____ When electrons in a covalent bond are shared unequally due to differences in electronegativity.

36. _____ Part of individuals will survive and reproduce. Those who survive are best adapted to the environment.

37. _____ A genetic change in a population of organisms which leads to a change from simple to complex