

Balance the following chemical reactions

1.  $1 \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ PO}_4$
2.  $\text{ } \underline{\hspace{1cm}} \text{ Ca} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Ca(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
3.  $\text{ } \underline{\hspace{1cm}} \text{ Fe}_2\text{SO}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Fe(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{SO}_4$
4.  $\text{ } \underline{\hspace{1cm}} \text{ C}_2\text{H}_6 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
5.  $\text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ PCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ PCl}_5$
6.  $\text{ } \underline{\hspace{1cm}} \text{ P}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ PH}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
7.  $2 \text{ AuCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ Au}_2\text{O}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
8.  $\text{ } \underline{\hspace{1cm}} \text{ Ca}_3\text{PO}_4 + \text{ } \underline{\hspace{1cm}} \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ Ca}_2\text{Al}_2\text{PO}_7$
9.  $2 \text{ Br}_2 + \text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ BrCl}$
10.  $\text{ } \underline{\hspace{1cm}} \text{ Fe} + \text{ } \underline{\hspace{1cm}} \text{ NaOH} + \text{ } \underline{\hspace{1cm}} \text{ Fe}_2\text{NaO}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
11.  $8 \text{ Fe} + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ Fe}$
12.  $2 \text{ CH}_3\text{OH} + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
13.  $\text{ } \underline{\hspace{1cm}} \text{ PCl}_5 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_3\text{PO}_4$
14.  $\text{ } \underline{\hspace{1cm}} \text{ PCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_3\text{PO}_3 + \text{ } \underline{\hspace{1cm}} \text{ HCl}$
15.  $\text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Fe(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ HCl}$
16.  $2 \text{ Mg} + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ MgO} + \text{ } \underline{\hspace{1cm}} \text{ Fe}$
17.  $\text{ } \underline{\hspace{1cm}} \text{ CaCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ Na}_2\text{CO}_3 + \text{ } \underline{\hspace{1cm}} \text{ CaCO}_3 + \text{ } \underline{\hspace{1cm}} \text{ NaCl}$
18.  $2 \text{ C}_2\text{H}_6 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
19.  $\text{ } \underline{\hspace{1cm}} \text{ H}_2\text{S} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Al(OH)}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
20.  $\text{ } \underline{\hspace{1cm}} \text{ Cl}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ HClO}_2$
21.  $\text{ } \underline{\hspace{1cm}} \text{ MnCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ HCl} + \text{ } \underline{\hspace{1cm}} \text{ MnO}_2 + \text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$