+WS 5.4	Ideal Gas L	.aw	(pg 1)		PV=nRT	R = 0.0821 L-atm
1. What v	olume would	3.00 m	noles of neon gas	have at 295 K and		mol-K
2. What v	olume would	4.3 mc	oles of hydrogen	gas occupy at 45°(C and 326.2 kP	Ans:
3. How m	uch pressure	would	4.85 moles of He	gas exert in a 4.50	0 L tank at 55°(Ans:
4. How m	any moles o	f CO ₂ o	could fit in a 475 n	nL bag at -22°C an	d 855 mmHg?	Ans:
5. How m	any grams of	oxyger	n gas are there in	a 2.3 L tank at 7.5	atm and 24°C	Ans:
6. How m	any molecule	es of N ₂	could fitin a 2.00	L soda bottle at 2	23°C and 755 m	Ans: mHg?
7. What p	ressure wou	ld be n	eeded to fit 35.0 (g of N₂ gas intoa 1	195 mL flask at	Ans: 0°C?
			of gas fit in a box re be (in °C)? (14		00 dm x 2.40 dn	Ans: n x 5.83 dm at 1.00
				old precisely 40.0 (remember it's a cub		
10. What		ld be o	ocupied by 16.0 (g of CH ₄ at 0°C an	od 760 mmHg?	Ans:
						Ans:

(continued on next page!)