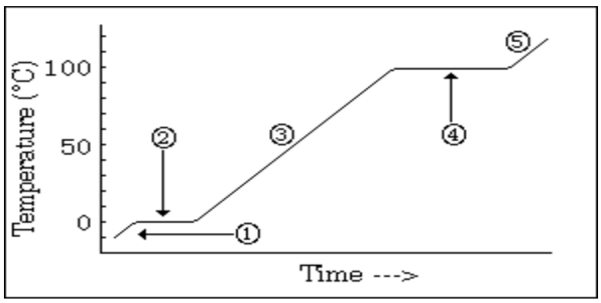
Heating Curve/Calorimetry Worksheet **Name**:_____
Calorimetry **Class**:_____ **Date**:_____



1) What is happening to the average kinetic energy of the molecules in the sample during section 2?_____

2) As a substance goes through section (2), what happens to the distance between the particles?_____

3) What is the name of the process happening during section (4)?_____

4) What would be the name of the process happening during section (4) if time were going the other way?_____

5) What is the melting point of this substance?_____

6) At what temperature would this sample finish boiling?_____

7) When this substance is melting, the temperature of the ice-water mixture remains constant because:
a. Heat is not being absorbed
b. The ice is colder that the water
c. Heat energy is being converted to potential energy
d. Heat energy is being converted to kinetic energy

8) When a given quantity of water is heated at a constant rate, the phase change from liquid to gas takes longer than the phase change from solid to liquid because
a. The heat of vaporization is greater than the heat of fusion
b. The heat of fusion is greater than the heat of vaporization
c. The average kinetic energy of the molecules is greater in steam than in water
d. Ice absorbs energy more rapidly than water does