

Name _____

Date _____ Pg _____

Chemistry Unit 7

Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation: _____

Suppose that 4 molecules of hydrogen gas and 4 molecules of oxygen gas react to form water.

Make a drawing that represents the reaction container before and after the reaction.



Before



After

_____ How many molecules of water can be produced?

_____ Which reactant is in excess? Why?

_____ How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reaction mixture:

Bal. Equation: _____

Before: _____

Change: _____

After: _____

According to the table you just made,

_____ How many molecules of water can be produced?

_____ Which reactant is in excess? Why?

_____ How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?