

t-Butyl Chloride - Concentration vs Rate of Solvolysis

Concentration of $t\text{-BuCl}$ (M)

Exp.	$t\text{-BuCl}$ (M)	Initial $t\text{-BuCl}$ (M)	$t\text{-BuCl}$ (M)	$t\text{-BuCl}$ (M)
1	0.00	0.00	0.00	0.00
2	0.00	0.00	0.00	0.00
3	0.00	0.00	0.00	0.00

Time (min) = 40.00

Exp.	$t\text{-BuCl}$ (M)	$t\text{-BuCl}$ reacted at time t (M)	Rate = $\frac{d[\text{t-BuCl}]}{dt}$	Initial $t\text{-BuCl}$ (M)
1	0.00			0.00
2	0.00			0.00
3	0.00			0.00

Time (min) = 100.00

Exp.	$t\text{-BuCl}$ (M)	$t\text{-BuCl}$ reacted at time t (M)	Rate = $\frac{d[\text{t-BuCl}]}{dt}$	Initial $t\text{-BuCl}$ (M)
1	0.00			0.00
2	0.00			0.00
3	0.00	0.00		0.00

Time (min) = 170.00

Exp.	$t\text{-BuCl}$ (M)	$t\text{-BuCl}$ reacted at time t (M)	Rate = $\frac{d[\text{t-BuCl}]}{dt}$	Initial $t\text{-BuCl}$ (M)
1	0.00			0.00
2	0.00	0.00		0.00
3	0.00	0.00		0.00

Initial $t\text{-BuCl}$ (M)	Initial $t\text{-BuCl}$ (M)	Initial $t\text{-BuCl}$ (M)	Initial $t\text{-BuCl}$ (M)	Initial $t\text{-BuCl}$ (M)	Rate	Concentration
0.00	0.00	0.00	0.00	0.00	0.00	
0.00	0.00	0.00	0.00	0.00	0.00	
0.00	0.00	0.00	0.00	0.00	0.00	

