Chemistry: Compounds & Formulas **TEST REVIEW**

1. Write Formulas for these compounds:

a. magnesium iodide _____ b. copper (I) oxide _____

c. potassium chlorate _____ d. iron (II) nitrite _____

e. lithium bisulfate _____ f. calcium hydroxide _____

g. silver chloride _____ h. mercury (I) phosphate _____

i. silicon tetraiodide _____ j. mercury (II) nitrate _____

2. Write the names of these compounds:

a. MnO_2 _____ b. $LiHCO_3$ _____

c. Cu_2SO_4 _____ d. P_3O_5 _____

e. N_2O _____ f. CuS _____

g. $FeSO_4$ _____ h. $(NH_4)_2Cr_2O_7$ _____

i. $Mg_3(PO_4)_2$ _____ j. Hg_2I_2 _____

k. $Zn(C_2H_3O_2)_2$ _____ l. Rb_2O _____

m. $Al_2(CrO_4)_3$ _____ n. $Al(NO_3)_3$ _____

3. Calculate the formula weights of these compounds to the nearest tenth:

a. H_2SO_4 _____ b. $MgCrO_4$ _____

c. $MgCl_2$ _____ d. $Sb(NO_3)_2$ _____

4. Calculate the percent composition of $Sr(HSO_3)_2$ to the nearest tenth percent.

a. %Sr _____ b. %H_____ c. % S _____ d. % O _____

5. Calculate the percent composition of SO_3 in H_2SO_4. % SO_3 _____

6. How many grams are in:

a. 1 mole $Al(SO_4)_2$ _____

b. 3.5 moles $HC_2H_3O_2$ _____