

Chemical Nomenclature – Ionic Compounds

Write the names and charges of the following ions. Most transition metals have more than one charge, but some have only one charge (Ag, Cd, Zn). See p. 121.

Li _____	Na _____	K _____
Be _____	Mg _____	Ca _____
Cu ⁺² _____	Zn _____	Ag _____
N _____	O _____	F _____
P _____	S _____	Br _____

Refer to Table 5.3 to answer the following questions:

Write a balanced formula for tin (II) oxide. _____

Write a balanced formula for tin (IV) oxide. _____

Name the following compound: MnO₂ _____

Name the following compound: Mn₂O₅ _____

Name the following binary ionic compounds:

a. Cs₂O _____

b. BaCl₂ _____

c. AgBr _____

d. ZnS _____

e. CuS _____

f. FeCl₃ _____

g. SnO₂ _____



Write formulas for the following ionic compounds:

a. chromium (III) oxide _____

b. aluminum oxide _____

c. lead (IV) oxide _____

d. manganese (VII) oxide _____

e. mercury (II) sulfide _____