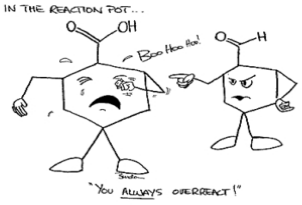
CHAPTER 6 REVIEW WORKSHEET

Name:_____ Block:_____ Date:_____

1. What type of chemical reaction involves two smaller molecules reacting to produce one large molecule?
   A. Decomposition
   B. Synthesis
   C. Single Replacement
   D. Combustion

2. Given the incomplete equation of a chemical reaction: $C_9H_6O_4 + O_2 \rightarrow$
Which of the following are products formed from this reaction?



I. H_2
II. CO_2
III. H_2O

   A. I and II only
   B. I and III only
   C. II and III only
   D. I, II and III

3. Which of the following represents a single replacement reaction?

I. $Sn + 2AgNO_3 \rightarrow Sn(NO_3)_2 + 2Ag$
II. Gold (II) cyanide + zinc → gold + zinc cyanide
III. Magnesium iodide reacts with bromine gas to produce magnesium bromide and iodine

   A. I and II only
   B. I and III only
   C. II and III only
   D. I, II and III

4. Sodium nitrate is produced as a result of mixing a solution of cadmium(II) nitrate with a solution of sodium sulphide. What is the other compound formed from this reaction?
   A. CdS
   B. $CdSO_4$
   C. NaS_2
   D. $CdNO_4$

5. What type of reaction is represented by the following word equation:

Potassium chlorate → Potassium chloride + oxygen

   A. Synthesis
   B. Decomposition
   C. Single replacement
   D. Double replacement

6. Predict the products of the following chemical reactions, classify the reaction and balance it.

Reactants		Products	Reaction Type
$2NaI + F_2$	→	$2NaF + I_2$	Single Replacement
$Cl_2 + 2KI$	→	$2KCl + I_2$	Single Replacement
$3AgNO_3 + Na_3PO_4$	→	$Ag_3PO_4 + 3NaNO_3$	Double Replacement
$2CH_3OH + O_2$	→	$2CO_2 + H_2O$	Combustion
$MgCl_2$	→	$Mg + Cl_2$	Decomposition
$3Sr(OH)_2 + 2H_3PO_4$	→	$Sr_3(PO_4)_2 + 6H_2O$	Neutralization