

Electron configuration practice

Name: _____

1. Write the orbital notation for the following:
 - a. carbon
 - b. neon
 - c. sulfur

2. Write the superscript notation for the following:
 - a. carbon
 - b. neon
 - c. sulfur

3. Write the noble gas notation for the following:
 - a. carbon
 - b. neon
 - c. sulfur

4. Write orbital notations for the following:
 - a. P
 - b. B
 - c. Na

5. Write superscript notation for atoms containing the following number of electrons:
 - a. 3
 - b. 6
 - c. 8
 - d. 13

6. Write noble gas notation for each of the following elements:
 - a. Cl
 - b. Ca
 - c. Se
 - d. Na
 - e. Sr
 - f. P

7. Write superscript notations for the following:
 - a. Na
 - b. Sr
 - c. P

8. Identify the following elements based on their electron configurations:
 - a. $1s^2 2s^2 2p^6 3s^2 3p^3$
 - b. $[\text{Ar}]4s^1$
 - c. $1s^2 2s^2 2p^1$
 - d. $1s^2 2s^2 2p^5$
 - e. $[\text{Ne}]3s^2$
 - f. $[\text{Ne}]3s^2 3p^5$
 - g. $[\text{Ar}]3d^6 4s^2$