

CHEMISTRY**ATOMIC STRUCTURE PRACTICE I**

X = element symbol

A = mass number [# of protons (p) + # neutrons (n)]

Z = atomic number [# of protons]

N = # of neutrons

A - Z = N

A typical isotopic symbol takes this form:



ex. The isotopic symbol for Fluorine would be



Fill in the missing items in the table below.

| Name | Symbol | Z | A | #p | #e | #n | Isotopic Symbol |
|-----------|--------|----|---|----|----|----|-----------------|
| | Na | | | | | | |
| | | 17 | | | | | |
| Potassium | | | | | | | |

Fill in the missing items in the table below.

| Name | Symbol | Z | A | #p | #e | #n | Isotopic Symbol |
|------|--------|---|---|----|----|----|-----------------|
| | P | | | | | | |
| Iron | | | | | | | |
| | | | | 53 | | | |

Fill in the missing items in the table below.

| Name | Symbol | Z | A | #p | #e | #n | Isotopic Symbol |
|--------|--------|----|---|----|----|----|-----------------|
| Silver | | | | | | | |
| | | 36 | | | | | |
| | W | | | | | | |