

## Writing Ionic Formulas



Salt = NaCl

Reminder:



Called the

\* If there is no number after the + or - sign, it implies a 1.

To determine the formula for an ionic compound,  
we use the CRISSCROSS RULE.

↓  
the oxidation # of the cation becomes the subscript for the anion &  
the oxidation # of the anion becomes the subscript for the cation

The + & - signs DO NOT crossover!



We don't write a subscript for the number 1)



We can reduce these subscripts because they are both divisible by 2. The formula becomes

Always reduce the formula numbers by the lowest common denominator. (Only if ALL the elements are divisible.)