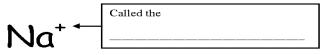
Writing Ionic Formulas



Reminder:



* If there is no number after the + or - sign, it implies a 1.

To determine the formula for an ionic compound, we use the <u>CRISSCROSS RULE</u>.

the oxidation # of the cation becomes the subscript for the anion & the oxidation # of the anion becomes the subscript for the cation

The + & - signs <u>DO NOT</u> crossover!

1. K⁺ S²⁻ becomes _____

We don't write a subscript for the number 1)

2. Ca^{2+} , S^{2-} becomes _____

We can reduce these subscripts because they are both divisible by 2. The formula becomes

Always reduce the formula numbers by the lowest common denominator. (Only if ALL the elements are divisible.)