

## Naming Ionic Compounds and Writing Ionic Formulas

Write the answers in your spiral notebooks.

1. Write the formulas for the following ions:

Ammonium	Sulfate	Potassium Ion
Tin (II)	Calcium Ion	Lead (IV)
Chromate	Nitride	Silver
Sulfide	Carbonate	Chlorate
Cyanide	Phosphate	Acetate
Iron (III)	Hydroxide	Copper (I)
Permanganate	Nitrate	Phosphide
Nitrite	Sulfite	Dichromate

2. Write the formulas for the following ionic compounds:

Barium Hydroxide	Beryllium Nitride	Potassium Bromide
Aluminum Chloride	Vanadium (III) Oxide	Lead (IV) Chlorate
Magnesium Sulfide	Silver Nitrite	Cadmium Phosphide
Ammonium Nitrate	Zinc Chlorite	Cesium Sulfate
Rubidium Phosphate	Chromium (II) Perchlorate	Ammonium Hydroxide
Iron (III) Acetate	Aluminum Nitride	Copper (II) Fluoride
Calcium Oxide	Manganese (III) Cyanide	Barium Sulfate
Sodium Carbonate	Tin (II) Sulfite	Strontium Dichromate

3. Write the names of the following ionic compounds:

NaCl	SnCl <sub>4</sub>	SnCl <sub>2</sub>
BaCl <sub>2</sub>	MgCO <sub>3</sub>	Co <sub>2</sub> O <sub>3</sub>
AlF <sub>3</sub>	Fe(ClO <sub>4</sub> ) <sub>3</sub>	Al <sub>2</sub> P <sub>3</sub>
FeO	Be(CH <sub>3</sub> COO) <sub>2</sub>	Ca(NO <sub>3</sub> ) <sub>2</sub>
KI	SnS <sub>2</sub>	K <sub>2</sub> SO <sub>3</sub>
AgCl	Mn(HCO <sub>3</sub> ) <sub>3</sub>	Fe <sub>2</sub> O <sub>3</sub>
Cu <sub>2</sub> O	Na <sub>3</sub> PO <sub>4</sub>	Ba(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>2</sub>
CaS	CoO	NaClO
(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>	CrCl <sub>3</sub>	PbO <sub>2</sub>