

## binary ionic compounds (Homework)

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Be sure to mention the file name: **WebAssign\_Chemistry\_13**

1.  
A simple ion with a +1 charge (for example,  $\text{Na}^+$ ) results when an atom \_\_\_\_\_ electrons.
2.  
Positive ions are called \_\_\_\_\_.  
whereas negative ions are called \_\_\_\_\_.
3.  
Simple negative ions formed from single atoms are given names that end with the letters \_\_\_\_\_.
4.  
How many electrons are contained in each of the following ions?  

(a) $\text{K}^+$	(b) $\text{Mn}^{2+}$	(c) $\text{Co}^{3+}$	(d) $\text{Co}^{2+}$
(e) $\text{Cr}^{3+}$	(f) $\text{I}^-$	(g) $\text{Fe}^{3+}$	(h) $\text{P}^{3-}$
5.  
For the following pairs of ions, use the concept that a chemical compound must have a net charge of zero to predict the formula of the simplest compound that the ions are most likely to form. (Type your answer using the format  $\text{CO}_2$  for  $\text{CO}_2$ .)  

(a) $\text{Na}^+$ and $\text{C}^4-$	(b) $\text{Sn}^{4+}$ and $\text{N}^{3-}$	(c) $\text{Fe}^{3+}$ and $\text{P}^{3-}$	(d) $\text{Sn}^{4+}$ and $\text{C}^{4-}$
(e) $\text{Mg}^{2+}$ and $\text{N}^{3-}$	(f) $\text{Mg}^{2+}$ and $\text{C}^{4-}$	(g) $\text{Fe}^{3+}$ and $\text{C}^{4-}$	(h) $\text{Fe}^{3+}$ and $\text{S}^{2-}$
6.  
Write the correct formula for the ionic compound composed of potassium and iodide. (Type your answer using the format  $\text{CH}_4$  for  $\text{CH}_4$ .)
- 7.