

## binary ionic compounds (Homework)

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1. A simple ion with a +1 charge (for example,  $\text{Na}^+$ ) results when an atom \_\_\_\_\_ electrons.
2. Positive ions are called \_\_\_\_\_.  
whereas negative ions are called \_\_\_\_\_.
3. Simple negative ions formed from single atoms are given names that end with the letters \_\_\_\_\_.
4. How many electrons are contained in each of the following ions?  
(a)  $\text{K}^+$       (b)  $\text{Mn}^{2+}$       (c)  $\text{Co}^{3+}$       (d)  $\text{Co}^{2+}$

(e)  $\text{Cr}^{3+}$

(f)  $\text{I}^-$

(g)  $\text{Fe}^{3+}$

$\text{MnO}_2$

$\text{NiF}_3$

$\text{CuS}$

$\text{CuBr}$

$\text{Fe}_2\text{O}_3$

$\text{SnF}_4$

$\text{Ga}_2\text{S}_3$

$\text{AgI}$

$\text{PbO}$

$\text{Pb}_2\text{S}_3$

$\text{Hg}_2\text{F}_2$

$\text{HgI}_2$

$\text{CdSe}$

$\text{ZnBr}_2$

$\text{ScCl}_3$

$\text{MnCl}_2$

$\text{NiCl}_2$