

Types of Chemical Reactions

Name _____

Period _____ Date _____

Directions: Label the following reactions as synthesis (S), decomposition (D), single replacement (SR), double replacement (DR) and combustion (C), then balance. Place a "B" at the end that are balanced as written.

1. $\text{NaCl} \rightarrow \text{Na} + \text{Cl}_2$ _____
2. $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$ _____
3. $\text{C}_2\text{H}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ _____
4. $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$ _____
5. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ _____
6. $\text{C}_8\text{H}_{18} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ _____
7. $\text{NaOH} + \text{HCl} \rightarrow \text{HOH} + \text{NaCl}$ _____
8. $\text{Na} + \text{HCl} \rightarrow \text{H}_2 + \text{NaCl}$ _____
9. $\text{K} + \text{Cl}_2 \rightarrow \text{KCl}$ _____
10. $\text{K} + \text{AgCl} \rightarrow \text{Ag} + \text{KCl}$ _____
11. $\text{C}_6\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ _____
12. $\text{Ca} + \text{S} \rightarrow \text{CaS}$ _____
13. $\text{KOH} + \text{HNO}_3 \rightarrow \text{HOH} + \text{KNO}_3$ _____
14. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ _____
15. $\text{HgO} \rightarrow \text{Hg} + \text{O}_2$ _____
16. $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$ _____
17. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ _____
18. $\text{C}_6\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ _____
19. $\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + \text{KNO}_3$ _____
20. $\text{H}_2 + \text{N}_2 \rightarrow \text{NH}_3$ _____