

### Nomenclature Worksheet #4 – Polyatomic Compounds

WRITE FORMULAS FOR EACH OF:

- |                                 |   |
|---------------------------------|---|
| 1. lithium nitrate _____        | 17. aluminum sulfate _____              |
| 2. sodium chlorate _____        | 18. aluminum phosphate _____            |
| 3. calcium nitrate _____        | 19. copper (I) sulfate _____            |
| 4. sodium phosphate _____       | 20. copper (I) phosphate _____          |
| 5. aluminum nitrate _____       | 21. zinc sulfate _____                  |
| 6. potassium sulfate _____      | 22. cobalt (III) chloride _____         |
| 7. ammonium nitrate _____       | 23. calcium hydroxide _____             |
| 8. ammonium oxide _____         | 24. sodium hydroxide _____              |
| 9. potassium chlorate _____     | 25. aluminum hydroxide _____            |
| 10. barium carbonate _____      | 26. ammonium hydroxide _____            |
| 11. iron (II) sulfate _____     | 27. sodium bicarbonate _____            |
| 12. iron (II) nitrate _____     | 28. ammonium bicarbonate _____          |
| 13. iron (II) phosphate _____   | 29. copper (II) bicarbonate _____       |
| 14. titanium (IV) sulfate _____ | 30. aluminum bicarbonate _____          |
| 15. nickel (III) nitrate _____  | 31. iron (III) hydrogen carbonate _____ |
| 16. silver nitrate _____        | 32. sulfate radical _____               |

NAME EACH OF THE FOLLOWING:

33.  $\text{CaSO}_4$  \_\_\_\_\_
34.  $\text{LiOH}$  \_\_\_\_\_
35.  $\text{Ba}(\text{OH})_2$  \_\_\_\_\_
36.  $\text{KNO}_3$  \_\_\_\_\_
37.  $\text{Pb}(\text{OH})_2$  \_\_\_\_\_
38.  $\text{Ba}_3(\text{PO}_4)_2$  \_\_\_\_\_
39.  $(\text{NH}_4)_2\text{SO}_4$  \_\_\_\_\_
40.  $\text{Mg}(\text{NO}_3)_2$  \_\_\_\_\_
41.  $\text{Ni}_2(\text{SO}_4)_3$  \_\_\_\_\_
42.  $\text{Pb}(\text{ClO}_3)_2$  \_\_\_\_\_
43.  $\text{Na}_2\text{SO}_4$  \_\_\_\_\_
44.  $\text{SrCO}_3$  \_\_\_\_\_