Name_____Per____

CHEMICAL REACTIONS WORKSHEET I

1. $Zn + 2HCl \text{ ----> } ZnCl_2 + H_2$ Single Replacement

2. $2H_2O \text{ ----> } 2H_2 + O_2$ Decomposition

3. $2NaOH + Pb(NO_3)_2 \text{ ----> } 2NaNO_3 + Pb(OH)_2$
Double Replacement

4. $H_2O + CO_2 \text{ ----> } H_2CO_3$ Already balanced, Synthesis

5. $2C_4H_{10} + 13O_2 \text{ ----> } 8CO_2 + 10H_2O$ Combustion

6. $CuSO_4 + Mg \text{ ----> } MgSO_4 + Cu$ Already balanced, Single Replacement

7. $2C_6H_6(l) + 15O_2(g) \text{ ----> } 12CO_2(g) + 6H_2O(l)$ Combustion

8. $H_2O(l) + SO_2(g) \text{ ----> } H_2SO_3(aq)$ Already balanced, Synthesis

9. $Na_2CO_3(aq) + 2NH_4OH(aq) \text{ ---> } (NH_4)_2CO_3(aq) + 2NaOH(aq)$ Double Replacement

10. $(NH_4)_2CO_3(aq) \text{ ----> } 2NH_3(g) + H_2O(l) + CO_2(g)$ Decomposition

11. $4HNO_3(aq) + Cu(s) \text{ ---> } Cu(NO_3)_2(aq) + 2NO_2(g) + 2H_2O(l)$ Synthesis (sort of)

12. $Pb(NO_3)_2(aq) + 2NaCl(aq) \text{ --> } PbCl_2(s) + 2NaNO_3(aq)$ Double Replacement

13. $2Fe(s) + O_2(g) \text{ --> } 2FeO(s)$ Synthesis, Combustion

14. $4Fe(s) + 3O_2(g) \text{ --> } 2Fe_2O_3(s)$ Synthesis, Combustion

15. $4FeO(s) + O_2(g) \text{ --> } 2Fe_2O_3(s)$ Synthesis, Combustion

16. $Au_2S_3(s) + 3H_2(g) \text{ --> } 2Au(s) + 3H_2S(g)$ Single Replacement