Classifying Chemical Reactions Worksheet

Name: _____ **Period:** _____

Classify each reaction as synthesis, decomposition, single replacement, double replacement or combustion. The equations are **not** balanced.

1. $PbCl_2 + AgNO_3 \rightarrow Pb(NO_3)_2 + AgCl$ _____
2. $NH_3 + HCl \rightarrow NH_4Cl$ _____
3. $AlCl_3 + Na_2SO_4 \rightarrow Al_2(SO_4)_3 + NaCl$ _____
4. $Zn + S \rightarrow ZnS$ _____
5. $Al_2(SO_4)_3 + BaCl_2 \rightarrow BaSO_4 + AlCl_3$ _____
6. $Al_2S_3 \rightarrow Al + S$ _____
7. $H_2SO_4 + Fe \rightarrow H_2 + FeSO_4$ _____
8. $C_{12}H_{22}O_{11} + O_2 \rightarrow CO_2 + H_2O$ _____
9. $Mg(OH)_2 + H_2SO_4 \rightarrow MgSO_4 + H_2O$ _____
10. $NaOH + CuSO_4 \rightarrow Na_2SO_4 + Cu(OH)_2$ _____
11. $C_4H_{12} + O_2 \rightarrow H_2O + CO_2$ _____
12. $Fe + O_2 \rightarrow Fe_2O_3$ _____
13. $Mg_3(PO_4)_2 + H_2 \rightarrow Mg + H_3PO_4$ _____
14. $NH_4NO_3 \rightarrow N_2O + H_2O$ _____
15. $Cl_2 + KBr \rightarrow KCl + Br_2$ _____