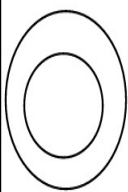


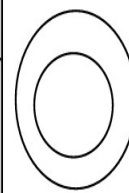
Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

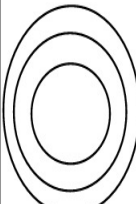

## BOHR ATOMIC MODELS

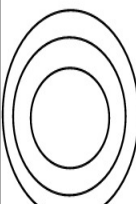
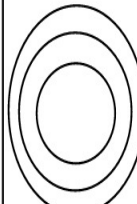
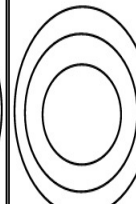
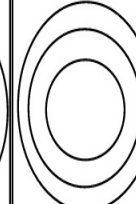
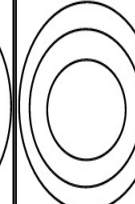
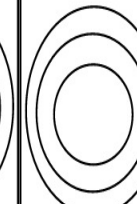
Hydrogen
Symbol _____
Atomic Number _____
Mass Number _____


**Procedure:**

1. Draw Bohr atomic models for each of the atoms using your Periodic Table
2. To represent the # of protons write a P- followed by the number of protons. Place in nucleus.
3. To represent the # of neutrons write a N- followed by the number of neutrons. Place in nucleus.
4. Use periodic table to determine how many electrons are in each orbital.
5. Use dots to represent the electrons. Pair electrons after the 1st orbital to make for easier counting.
6. Be sure to write the symbol, atomic #, and mass # for each element.
7. See Carbon as an example of what your Bohr model should look like.
8. Answer "Atomic Models Questions" after you have finished.

Helium
Symbol _____
Atomic Number _____
Mass Number _____


Lithium	Beryllium
Symbol _____	Symbol _____
Atomic Number _____	Atomic Number _____
Mass Number _____	Mass Number _____
	

Boron	Carbon	Nitrogen	Oxygen	Fluorine	Neon
Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____
Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____
Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____
					

Sodium	Magnesium	Aluminum	Silicon	Phosphorus	Sulfur	Chlorine	Argon
Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____	Symbol _____
Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____	Atomic Number _____
Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____	Mass Number _____
