

CHEMISTRY - CHAPTER 3: EQUATIONS

BALANCING EQUATIONS: FORMULAS GIVEN

Practice Sheet 10

Balance the following equations:

1. $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
2. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
3. $\text{CaCl}_2 + \text{Na}_2\text{CO}_3 \rightarrow \text{CaCO}_3$
4. $\text{MgSO}_4 + \text{NaCl} \rightarrow \text{Na}_2\text{SO}_4$
5. $\text{Ba} + \text{O} \rightarrow \text{Ba}_2\text{O} + \text{BaO} + \text{O}_2$
6. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO} + \text{MgO}_2$
7. $\text{BaO} + \text{BaO}_2 \rightarrow \text{Ba}_2\text{O}$
8. $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2$
9. $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow \text{Fe}_3\text{O}_4 + \text{CO}_2$
10. $\text{MgCl}_2 + \text{NaOH} \rightarrow \text{Mg}(\text{OH})_2 + \text{NaCl}$
11. $\text{AgNO}_3 + \text{CaCl}_2 \rightarrow \text{AgCl} + \text{Ca}(\text{NO}_3)_2$
12. $\text{Ba} + \text{O}_2 \rightarrow \text{BaO} + \text{BaO}_2$
13. $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2$
14. $\text{BaCl}_2 + \text{Mg}(\text{NO}_3)_2 \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{MgCl}_2$
15. $\text{CaCl}_2 + \text{O}_2 \rightarrow \text{CaO} + \text{CaO}_2$
16. $\text{Fe} + \text{BaO} \rightarrow \text{BaFeO} + \text{Fe}$
17. $\text{Fe} + \text{BaO}_2 \rightarrow \text{BaFeO}_2$
18. $\text{Ba}(\text{NO}_3)_2 + \text{KNO}_3 \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{BaO}$
19. $\text{MgCl}_2 + \text{KNO}_3 \rightarrow \text{MgCl}_2 + \text{BaO}$
20. $\text{CaCl}_2 + \text{Ca}(\text{NO}_3)_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{CaCl}_2 + \text{BaO}$
21. $\text{Mg}(\text{NO}_3)_2 \rightarrow \text{Mg} + \text{CaO} + \text{BaO}$
22. $\text{CaCl}_2 + \text{BaO} \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{BaO}$
23. $\text{MgCl}_2 + \text{BaO} \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{CaO}$
24. $\text{Ca}(\text{NO}_3)_2 + \text{BaO} + \text{Fe}_2\text{O}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{BaO}$
25. $\text{Fe}_2\text{O}_3 + \text{BaO} \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{BaO}$
26. $\text{Fe} + \text{BaO} \rightarrow \text{Fe}_2\text{O}_3 + \text{BaO}$
27. $\text{BaO} + \text{BaO}_2 \rightarrow \text{Ba} + \text{O}_2 + \text{BaO}$
28. $\text{Ba}(\text{NO}_3)_2 + \text{O}_2 + \text{Fe}_2\text{O}_3 \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{BaO}$
29. $\text{Ba}(\text{NO}_3)_2 + \text{BaO}_2 + \text{BaO} \rightarrow \text{Ba}(\text{NO}_3)_2 + \text{BaO} + \text{Ca}(\text{NO}_3)_2 + \text{BaO}$