

Chemist: Key

**Counting Atoms Worksheet # 1**

1. Determine the subatomic particles in each of the following::

	p <sup>+</sup>	e <sup>-</sup>	n <sup>0</sup>	Atomic #	Mass #
<sup>27</sup> Al	13	13	14	13	27
Br	35	35	45	35	80
Fe	26	26	30	26	56
Ca <sup>+2</sup>	20	18	20	20	40
O <sup>-2</sup>	8	10	8	8	16

2. Complete the following table:

	p <sup>+</sup>	e <sup>-</sup>	n <sup>0</sup>	Atomic #	Mass #
K <sup>+1</sup>	19	18	21	19	40
Mg	12	12	12	12	24
Sr <sup>+2</sup>	38	36	53	38	91
<del>Al</del>	~~~~~				
F <sup>-1</sup>	9	10	10	9	19

3. Carbon consists of 98.89% <sup>12</sup>C (12.00000), and 1.110% <sup>13</sup>C (13.00335). Calculate the atomic weight of carbon to four significant figures.

$$(.9889)(12.00000) + (.01110)(13.00335) = 11.869 + .144337185 = 12.01$$

4. Gallium consists of two natural isotopes, <sup>69</sup>Ga (68.9257) makes up 60.40% of the total. Calculate the % abundance and the mass of the other isotope.

$$(.6040)(68.9257) + (.3960)(x) = 69.723$$

$$41.6311228 + .3960x = 69.723$$

$$x = 70.93908384 = 70.94$$