

Types of Chemical Reaction Worksheet

Name: _____

Balance the reactions and indicate which types of chemical reaction are being represented:

- a) $\text{NaBr} + \text{Ca(OH)}_2 \rightarrow \text{CaBr}_2 + \text{NaOH}$ Reaction Type : _____
- b) $\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$ Reaction Type : _____
- c) $\text{C}_5\text{H}_8\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ Reaction Type : _____
- d) $\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$ Reaction Type : _____
- e) $\text{Li}_3\text{N} + \text{NH}_4\text{NO}_3 \rightarrow \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$ Reaction Type : _____
- f) $\text{HBr} + \text{Al(OH)}_3 \rightarrow \text{H}_2\text{O} + \text{AlBr}_3$ Reaction Type : _____
- g) $\text{Na}_3\text{PO}_4 + \text{KOH} \rightarrow \text{NaOH} + \text{K}_3\text{PO}_4$ Reaction Type _____
- h) $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + \text{LiCl}$ Reaction Type _____
- i) $\text{C}_6\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ Reaction Type _____
- j) $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ Reaction Type _____
- k) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ Reaction Type _____
- l) $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_3$ Reaction Type _____
- m) $\text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + \text{RbF}$ Reaction Type _____
- n) $\text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + \text{Ag}$ Reaction Type _____
- o) $\text{C}_3\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ Reaction Type _____
- p) $\text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe(C}_5\text{H}_5)_2$ Reaction Type _____
- q) $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + \text{Cl}_2$ Reaction Type _____
- r) $\text{MgI}_2 + \text{Mn(SO}_3)_2 \rightarrow \text{MgSO}_3 + \text{MnI}_4$ Reaction Type _____
- s) $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ Reaction Type _____
- t) $\text{NO}_2 \rightarrow \text{O}_2 + \text{N}_2$ Reaction Type _____