Name: _____

Date: _____ Block: _____

Predicting Products of Chemical reactions

Directions: First, determine the type of reaction or single replacement (s) double replacement (r) synthesis (d) decomposition and (c) combustion. Secondly, predict the products in the reaction and write the correct chemical formula for each product. Finally, balance each reaction using the correct coefficients.

1. **Double replacement** $2\ AgNO_3 + CaCl_2 \rightarrow 2\ AgCl + Ca(NO_3)_2$

2. _____ $Al + Cl_2 \rightarrow$

3. _____ $FeCl_3 + KOH \rightarrow$

4. _____ $C_2H_2 + O_2 \rightarrow$

5. _____ $Al_2(SO_4)_3 + Pb(NO_3)_2 \rightarrow$

6. _____ $Ba + MgI_2 \rightarrow$

7. _____ $Ca(OH)_2 + heat \rightarrow$

8. _____ $H_2SO_4 + KOH \rightarrow$

9. _____ $Zn + HCl \rightarrow$
(Zinc is a 2+ ion.)

10. _____ $K + Br_2 \rightarrow$

11. _____ $Mg + FeSO_4 \rightarrow$

12. _____ $NaOH + HCl \rightarrow$

13. _____ $C_3H_{12} + O_2 \rightarrow$

14. _____ $H_2CO_3 \rightarrow$

15. _____ $Ca + O_2 \rightarrow$

16. _____ $2\ Li + Cu(NO_3)_2 \rightarrow$