Sig Fig Calculation and Nomenclature Worksheet

1) Do the indicated arithmetic and give the answer to the correct number of significant figures:

a) [illegible]	b) [illegible]
c) [illegible]	d) [illegible]
e) [illegible]	f) [illegible]
g) [illegible]	h) [illegible]
i) [illegible]	j) [illegible]
k) [illegible]	l)

2) **Provide names for the following compounds and give the general category they belong to (i.e. ionic, molecular, acid/base or hydrate):**

Na_2SO_3 sodium sulfite - ionic

$(NH_4)_3PO_4$ ammonium phosphate - ionic

$MnCl_2$ manganese (II) chloride - ionic

$Ca(OH)_2$ calcium hydroxide - base

H_2SO_4 (aq) sulfuric acid - acid

P_2O_5 diphosphorus pentoxide - molecular

HBr (aq) hydrobromic acid - acid

$Hg(NO_3)_2$ mercury(II) nitrate - ionic

K_2HPO_4 potassium hydrogen phosphate - ionic

CaH_2 calcium hydride - ionic

$NaHCO_3$ sodium bicarbonate - ionic

$Cu(ClO_3)_2$ copper(II) chlorate - ionic

$Pb(C_2O_4)_2$ (Latin) Plumbic Oxalate - ionic

$Al_2(SO_3)_3$ aluminum sulfite - ionic

$Fe(NO_3)_3$ iron(III) nitrate - ionic

NO_2 nitrogen dioxide - molecular

B_2O_3 diboron trioxide - molecular

$Fe(ClO_4)_3$ iron(III) perchlorate - ionic

HCN (aq) hydrocyanic acid - acid

$CuSO_4 \cdot 5H_2O$ copper(II) sulfate pentahydrate

Br_2 bromine - element

MgF_2 magnesium fluoride - ionic

PH_3 phosphorus trihydride - molecular

$Al(NO_2)_3$ aluminum nitrite - ionic

$CoCl_2 \cdot 6H_2O$ cobalt(II) chloride hexahydrate

Cu_2CrO_4 (Latin) cuprous chromate - ionic