

Balance the following chemical reactions

1. $1 \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ PO}_4$
2. $\text{ } \underline{\hspace{1cm}} \text{ Ca} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Ca(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
3. $\text{ } \underline{\hspace{1cm}} \text{ Fe}_2\text{SO}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Fe(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{SO}_4$
4. $\text{ } \underline{\hspace{1cm}} \text{ C}_2\text{H}_6 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
5. $\text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ PCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ PCl}_5$
6. $\text{ } \underline{\hspace{1cm}} \text{ P}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ PH}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
7. $2 \text{ AuCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ Au}_2\text{O}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
8. $\text{ } \underline{\hspace{1cm}} \text{ Ca}_3\text{PO}_4 + \text{ } \underline{\hspace{1cm}} \text{ AlPO}_4 + \text{ } \underline{\hspace{1cm}} \text{ Ca}_2\text{Al}_2\text{PO}_7$
9. $2 \text{ Br}_2 + \text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ BrCl}$
10. $\text{ } \underline{\hspace{1cm}} \text{ Fe} + \text{ } \underline{\hspace{1cm}} \text{ NaOH} + \text{ } \underline{\hspace{1cm}} \text{ Fe}_2\text{NaO}_4 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
11. $8 \text{ Fe} + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ Fe}$
12. $2 \text{ CH}_3\text{OH} + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
13. $\text{ } \underline{\hspace{1cm}} \text{ PCl}_5 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_3\text{PO}_4$
14. $\text{ } \underline{\hspace{1cm}} \text{ PCl}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_3\text{PO}_3 + \text{ } \underline{\hspace{1cm}} \text{ HCl}$
15. $\text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Fe(OH)}_2 + \text{ } \underline{\hspace{1cm}} \text{ HCl}$
16. $2 \text{ Mg} + \text{ } \underline{\hspace{1cm}} \text{ FeCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ MgO} + \text{ } \underline{\hspace{1cm}} \text{ Fe}$
17. $\text{ } \underline{\hspace{1cm}} \text{ CaCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ Na}_2\text{CO}_3 + \text{ } \underline{\hspace{1cm}} \text{ CaCO}_3 + \text{ } \underline{\hspace{1cm}} \text{ NaCl}$
18. $2 \text{ C}_2\text{H}_6 + \text{ } \underline{\hspace{1cm}} \text{ O}_2 + \text{ } \underline{\hspace{1cm}} \text{ CO}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$
19. $\text{ } \underline{\hspace{1cm}} \text{ H}_2\text{S} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ Al(OH)}_3 + \text{ } \underline{\hspace{1cm}} \text{ H}_2$
20. $\text{ } \underline{\hspace{1cm}} \text{ Cl}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O} + \text{ } \underline{\hspace{1cm}} \text{ HClO}_2$
21. $\text{ } \underline{\hspace{1cm}} \text{ MnCl}_2 + \text{ } \underline{\hspace{1cm}} \text{ HCl} + \text{ } \underline{\hspace{1cm}} \text{ MnO}_2 + \text{ } \underline{\hspace{1cm}} \text{ Cl}_2 + \text{ } \underline{\hspace{1cm}} \text{ H}_2\text{O}$