Name: _____ **Date:** _____

Strong Acids and Bases

1. Calculate the pH of a solution with $[H^+] = 5 \times 10^{-5}$ M.

   pH = 4.3

2. Calculate the pH of a solution with $[H^+] = 1$ M.

   pH = 0

3. Calculate the pH of a 0.01 M solution of HCl.

   pH = 2

4. Calculate the pH of a 0.05 M solution of NaOH.

   pH = 1.30

5. Calculate the pH of a 7.5×10^{-6} M solution of $Mg(OH)_2$.

   pH = 9.176

6. Find $[H^+]$ of a solution with pH = 3.

   $[H^+] = 10^{-3}$ M