

Ionic Chemical Formulas/PreAICE Chemistry

NAME: _____ PERIOD _____

Directions: Using the Chemical Formulas program sheet, fill in the table below. Use subscripts & superscripts. Remember to use roman numerals to show the valence of the transition metals in the compound name.

METAL CATION (+)	CATION FORMULA	NONMETAL ANION (-)	ANION FORMULA	COMPOUND FORMULA	COMPOUND NAME
sodium	Na+1	hydroxide	OH-	NaOH	sodium hydroxide
magnesium	Mg+2	phosphate	PO ₄ -3	Mg ₃ (PO ₄) ₂	magnesium phosphate
calcium	Ca+2	chloride	Cl-	CaCl ₂	calcium chloride
gold(I)	Au+1	sulfite	SO ₃ -2	Au ₂ SO ₃	gold(I) sulfite
copper(II)	Cu+2	hydroxide	OH-	Cu(OH) ₂	copper(II) hydroxide
Iron(III)	Fe+3	telluride	Te-2	Fe ₂ Te ₃	iron(III) telluride
magnesium	Mg+2	bromide	Br-1	MgBr ₂	magnesium bromide
manganese(II)	Mn+2	sulfide	S-2	MnS	manganese(II) sulfide
boron	B+3	oxide	O-2	B ₂ O ₃	boron oxide
potassium	K+	nitrate	NO ₃ -	KNO ₃	potassium nitrate
barium	Ba+2	bromide	Br-	BaBr ₂	barium bromide
strontium	Sr+2	sulfate	SO ₄ -2	SrSO ₄	strontium sulfate
iron(III)	Fe+3	oxide	O-2	Fe ₂ O ₃	iron(III) oxide
silver(I)	Ag+1	chloride	Cl-	AgCl	silver(I) chloride
aluminum	Al+3	carbonate	CO ₃ -2	Al ₂ (CO ₃) ₃	aluminum carbonate
gold(II)	Au+2	nitride	N-3	Au ₃ N ₂	gold(II) nitride
cesium	Cs+1	phosphate	PO ₄ -3	Cs ₃ PO ₄	cesium phosphate
ammonium	NH ₄ +	sulfide	S-2	(NH ₄) ₂ S	ammonium sulfide
zinc(II)	Zn+2	nitrite	NO ₂ -	Zn(NO ₂) ₂	zinc(II) nitrite
chromium(I)	Cr+1	nitride	N-3	Cr ₃ N	chromium(I) nitride
mercury(II)	Hg+2	oxide	O-2	HgO	mercury(II) oxide
sodium	Na+	sulfite	SO ₃ -2	Na ₂ SO ₃	sodium sulfite
magnesium	Mg+2	oxide	O-2	MgO	magnesium oxide