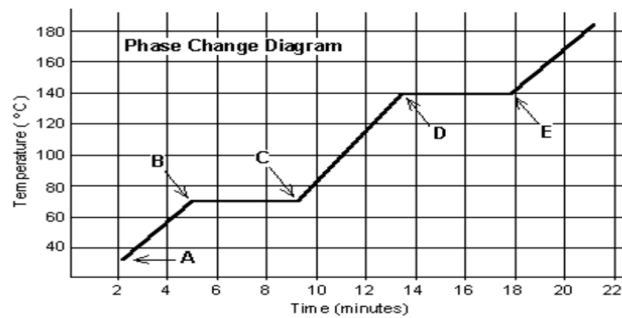
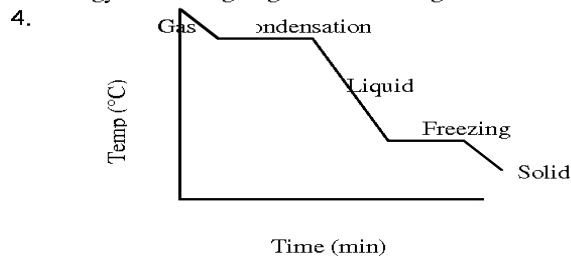


## Phase change review sheet answers



**A to B** = Solid  
**B to C** = melting point  
**C to D** = liquid  
**D to E** = boiling point, evaporation point, or vaporization point  
**E and higher** = gas

- 1.
2. this graph is endothermic
3. the temperature is constant from B to C and from D to E because the heat (energy) being absorbed is used to break the intermolecular forces (if this were a cooling curve the temperature would remain constant during a phase change because the energy would be going toward forming intermolecular forces)



This cooling curve is exothermic – energy (heat) is released to the surroundings in order to cool it down! (EXO – think “heat exits”)

5. The total energy absorbed in the warming curve is equal to the total energy released in the cooling curve - the amount of energy when you compare them is constant.
6. Temperature is a measure of the average kinetic energy of molecules. Thermal energy is heat energy in transfer. Heat is a form of energy. When heat is in the process of being transferred, it is called thermal energy. The higher the temperature, the faster the molecules being transferred will move.
- 7.