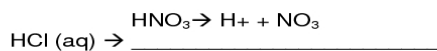


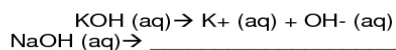
Worksheet: Acids, Bases, and Salts Review

Name _____

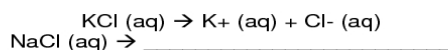
- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form _____ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form _____ ions.



- Salts are compounds that dissociate in aqueous solution releasing neither _____ ions nor _____ ions.



Using the Arrhenius definition, classify the following examples as acids, bases, or salts:

HBr _____	KCl _____
Mg(OH) ₂ _____	H ₃ PO ₄ _____
HCl _____	HClO _____
KNO ₂ _____	Al(OH) ₃ _____
HFO ₄ _____	KC ₂ H ₃ O ₂ _____
Ba(OH) ₂ _____	NaCl _____

Acids and bases can also be identified using an operational definition. Operational definitions are simply a list of properties.

ACIDS:

- ♦ A _____ taste is a characteristic property of all acids in aqueous solution.
- ♦ Acids react with some metals to produce _____ gas.
- ♦ Because aqueous acid solutions conduct electricity, they are identified as _____.
- ♦ Acids react with bases to produce a _____ and water.
- ♦ Acids turn _____ different colors.

BASES:

- ♦ Bases tend to taste _____ and feel _____.
- ♦ Like acids, aqueous basic solutions conduct _____, and are identified as _____.
- ♦ Bases react with _____ to produce a salt and _____.
- ♦ Bases turn _____ different colors.

Naming Acids

- Binary acids consist of _____ elements, the first being _____.
- Ternary acids consist of _____ elements. Do NOT use a prefix. _____-ate becomes _____ and _____-ite becomes _____.

1. Give the word equation for the neutralization reaction of an acid and a base.

2. Complete these equations by creating the formula and balancing:

Hydrochloric acid + lithium hydroxide →

Acetic acid + magnesium hydroxide →

Aluminum + carbonic acid →

Magnesium + phosphoric acid →