

- d. $\text{Mg}(\text{H}_2\text{PO}_4)_2 \rightarrow \text{Mg} + 2\text{H}_2 + 2\text{P} + 4\text{O}_2$
Reaction Type decomposition
- e. $\text{Cl}_2 + 2\text{KI} \rightarrow \text{I}_2 + 2\text{KCl}$
Reaction Type SR
- f. $2\text{Al} + 3\text{F}_2 \rightarrow 2\text{AlF}_3$
Reaction Type synthesis
- g. $\text{Zn} + \text{Cu}(\text{ClO}_3)_2 \rightarrow \text{Cu} + \text{Zn}(\text{ClO}_3)_2$
Reaction Type SR
- h. $2\text{Co}(\text{NO}_3)_3 + 3(\text{NH}_4)_2\text{C}_2\text{O}_4 \rightarrow \text{Co}_2(\text{C}_2\text{O}_4)_3 + 6\text{NH}_4\text{NO}_3$
Reaction Type DR
 $\begin{array}{l} \text{Co}^{3+} \times \text{NH}_4^+ \\ \text{NO}_3^- \times \text{C}_2\text{O}_4^{2-} \end{array}$
- i. $\text{C}_{12}\text{H}_{25}\text{OH} + 18\text{O}_2 \rightarrow 12\text{CO}_2 + 13\text{H}_2\text{O}$
Reaction Type Combustion
- j. $\text{Sr}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow 2\text{H}_2\text{O} + \text{Sr}(\text{NO}_3)_2$
Reaction Type neutralization or DR
- k. $4\text{V} + 5\text{O}_2 \rightarrow 2\text{V}_2\text{O}_5$ (or $2, \frac{5}{2}, 1$)
(Assume combining capacity of V is 5+)
Reaction Type Synthesis
- l. $\text{Rb}_3\text{AsO}_4 \rightarrow 3\text{Rb} + \text{As} + 2\text{O}_2$
Reaction Type decomposition
- m. $3\text{CsOH} + \text{H}_3\text{PO}_4 \rightarrow 3\text{H}_2\text{O} + \text{Cs}_3\text{PO}_4$
Reaction Type neutralization or DR
- n. $3\text{Ni}(\text{NO}_3)_2 + 2(\text{NH}_4)_3\text{PO}_4 \rightarrow \text{Ni}_3(\text{PO}_4)_2 + 6\text{NH}_4\text{NO}_3$
Reaction Type DR
 $\begin{array}{l} \text{Ni}^{2+} \times \text{NH}_4^+ \\ \text{NO}_3^- \times \text{PO}_4^{3-} \end{array}$