

Name _____ Period ____ Date ____

Molecular and Ionic Compounds
Practice Worksheet

1. Name the following molecular compounds

- | | |
|---------------------------------|-------------------------------|
| a. PCl_2 _____ | f. SiO_2 _____ |
| b. CCl_4 _____ | g. CO _____ |
| c. NO_2 _____ | h. CO_2 _____ |
| d. N_2F_2 _____ | i. H_2O _____ |
| e. P_4O_6 _____ | j. Cl_2 _____ |

2. Write formulas for the following molecular compounds

- | | |
|------------------------------|-----------------------------------|
| a. Water _____ | f. Sulfur dioxide _____ |
| b. Carbon dioxide _____ | g. Sulfur hexafluoride _____ |
| c. Silicon dioxide _____ | h. Nitrogen tribromide _____ |
| d. Dichlorine monoxide _____ | i. Dinitrogen tetrafluoride _____ |
| e. Nitrogen dioxide _____ | j. Diphosphorus pentoxide _____ |

3. Write the formulas for the following acids

- | | |
|----------------------------|--------------------------|
| a. Acetic acid _____ | d. Phosphoric acid _____ |
| b. Hydrochloric acid _____ | e. Carbonic acid _____ |
| c. Sulfuric acid _____ | f. Nitric acid _____ |

4. Name the following ionic compounds

- | | |
|-----------------------------------|-------------------------------|
| a. KBr _____ | e. CuI _____ |
| b. FeCl_3 _____ | f. K_2S _____ |
| c. NiCl_2 _____ | g. CaCO_3 _____ |
| d. Na_3PO_4 _____ | h. LiNO_3 _____ |

5. Name and indicate whether the compounds are ionic or molecular

- | |
|---------------------------------------|
| a. SeF_4 _____ |
| b. KBr _____ |
| c. H_2 _____ |
| d. CO_2 _____ |
| e. H_2CO_3 _____ |
| f. $\text{Cu}(\text{OH})_2$ _____ |
| g. NaHSO_4 _____ |
| h. $(\text{NH}_4)_2\text{SO}_4$ _____ |