

**Types of Chemical Reaction Worksheet**

Name: \_\_\_\_\_

*Balance the reactions and indicate which types of chemical reaction are being represented:*

- a) \_\_\_\_ NaBr + \_\_\_\_ Ca(OH)<sub>2</sub> → \_\_\_\_ CaBr<sub>2</sub> + \_\_\_\_ NaOH      Reaction Type : \_\_\_\_\_
- b) \_\_\_\_ NH<sub>3</sub> + \_\_\_\_ H<sub>2</sub>SO<sub>4</sub> → \_\_\_\_ (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>      Reaction Type : \_\_\_\_\_
- c) \_\_\_\_ C<sub>5</sub>H<sub>8</sub>O + \_\_\_\_ O<sub>2</sub> → \_\_\_\_ CO<sub>2</sub> + \_\_\_\_ H<sub>2</sub>O      Reaction Type : \_\_\_\_\_
- d) \_\_\_\_ Pb + \_\_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_\_ H<sub>2</sub> + \_\_\_\_ Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>      Reaction Type : \_\_\_\_\_
- e) \_\_\_\_ Li<sub>3</sub>N + \_\_\_\_ NH<sub>4</sub>NO<sub>3</sub> → \_\_\_\_ LiNO<sub>3</sub> + \_\_\_\_ (NH<sub>4</sub>)<sub>3</sub>N      Reaction Type : \_\_\_\_\_
- f) \_\_\_\_ HBr + \_\_\_\_ Al(OH)<sub>3</sub> → \_\_\_\_ H<sub>2</sub>O + \_\_\_\_ AlBr<sub>3</sub>      Reaction Type : \_\_\_\_\_
- g) \_\_\_\_ Na<sub>3</sub>PO<sub>4</sub> + \_\_\_\_ KOH → \_\_\_\_ NaOH + \_\_\_\_ K<sub>3</sub>PO<sub>4</sub>      Reaction Type \_\_\_\_\_
- h) \_\_\_\_ MgCl<sub>2</sub> + \_\_\_\_ Li<sub>2</sub>CO<sub>3</sub> → \_\_\_\_ MgCO<sub>3</sub> + \_\_\_\_ LiCl      Reaction Type \_\_\_\_\_
- i) \_\_\_\_ C<sub>6</sub>H<sub>12</sub> + \_\_\_\_ O<sub>2</sub> → \_\_\_\_ CO<sub>2</sub> + \_\_\_\_ H<sub>2</sub>O      Reaction Type \_\_\_\_\_
- j) Pb + FeSO<sub>4</sub> → PbSO<sub>4</sub> + Fe      Reaction Type \_\_\_\_\_
- k) CaCO<sub>3</sub> → CaO + CO<sub>2</sub>      Reaction Type \_\_\_\_\_
- l) P<sub>4</sub> + O<sub>2</sub> → P<sub>2</sub>O<sub>3</sub>      Reaction Type \_\_\_\_\_
- m) RbNO<sub>3</sub> + BeF<sub>2</sub> → Be(NO<sub>3</sub>)<sub>2</sub> + RbF      Reaction Type \_\_\_\_\_
- n) AgNO<sub>3</sub> + Cu → Cu(NO<sub>3</sub>)<sub>2</sub> + Ag      Reaction Type \_\_\_\_\_
- o) C<sub>3</sub>H<sub>6</sub>O + O<sub>2</sub> → CO<sub>2</sub> + H<sub>2</sub>O      Reaction Type \_\_\_\_\_
- p) C<sub>5</sub>H<sub>5</sub> + Fe → Fe(C<sub>5</sub>H<sub>5</sub>)<sub>2</sub>      Reaction Type \_\_\_\_\_
- q) SeCl<sub>6</sub> + O<sub>2</sub> → SeO<sub>2</sub> + Cl<sub>2</sub>      Reaction Type \_\_\_\_\_
- r) MgI<sub>2</sub> + Mn(SO<sub>3</sub>)<sub>2</sub> → MgSO<sub>3</sub> + MnI<sub>4</sub>      Reaction Type \_\_\_\_\_
- s) O<sub>3</sub> → O· + O<sub>2</sub>      Reaction Type \_\_\_\_\_
- t) NO<sub>2</sub> → O<sub>2</sub> + N<sub>2</sub>      Reaction Type \_\_\_\_\_