

Atomic Structure & Periodic Table Review – Ch. 10

ATOMIC STRUCTURE – COMPLETE THE FOLLOWING TABLE.

	Isotope Name	Atomic #	Mass #	p	n	e
1.	hydrogen-2					
2.		1	1			
3.				6	6	
4.			7	3		
5.	chlorine-35					
6.		19			20	
7.			24	12		
8.		33	74			
9.				47	62	
10.			107			47
11.		16			16	
12.	uranium-235					

AVERAGE ATOMIC MASS – CALCULATE THE AVG. ATOMIC MASS.

13. 80% ^{127}I , 17% ^{136}I , 3% ^{128}I 16. 99% ^1H , 0.8% ^2H , 0.2% ^3H
 14. 5 of 10 are ^{197}Au , 5 of 10 are ^{198}Au 17. 95% ^{14}N , 3% ^{15}N , 2% ^{16}N
 15. 3 of 20 are ^{55}Fe , 17 of 20 are ^{56}Fe 18. 9 of 10 are ^{12}C , 1 of 10 are ^{14}C

PERIODIC TRENDS

WHICH ATOM HAS THE LARGER RADIUS?

19. C vs. Sn 21. Li vs. K
 20. Sr vs. In 22. Ba vs. O

WHICH PARTICLE HAS THE LARGER RADIUS?

23. Mg vs. Mg^{2+} 25. Na vs. Na^+
 24. I vs. I^- 26. Se vs. Se^{2-}

WHICH ATOM HAS THE LARGER 1ST IONIZATION ENERGY?

27. Fr vs. Ne 29. Ar vs. Rn
 28. Na vs. Si 30. Be vs. Ca

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