

Mobility

(How quickly molecules move through a given substance)

What?

Diffusion is the movement of particles from an area of high concentration to an area of low concentration. It occurs in all states of matter. In a liquid, the particles are close together and can move past each other. In a gas, the particles are far apart and move quickly. In a solid, the particles are packed closely together and can only vibrate in place. Diffusion is faster in gases than in liquids, and faster in liquids than in solids.

Model 1 – Concentration Difference

Concentration Difference I



Concentration Difference II



1. Explain the difference between the two models. (2 points each)

- a. Model 1: The concentration difference is the only factor. The particles move from the high concentration area to the low concentration area.
- b. Model 2: The concentration difference is the only factor. The particles move from the high concentration area to the low concentration area.
- c. Model 3: The concentration difference is the only factor. The particles move from the high concentration area to the low concentration area.
- d. Model 4: The concentration difference is the only factor. The particles move from the high concentration area to the low concentration area.