Write the Ksp expression for the following:

1) $AlPO_4$
2) $Cu_3(PO_4)_2$
3) $CuSCN$
4) $AgCN$
5) $Zn_3(AsO_4)_2$
6) $Mn(IO_3)_2$

Given the following solubilities, calculate the solubility product (Ksp) to each compound.

7. $AgNO_3$ 1.55 g/L
8. $SrCO_3$ 7.5×10^{-5} M
9. BaC_2O_4 9.3×10^{-5} g/ml
10. CuCl 1.01×10^{-3} M

Given the following Ksp values, calculate the solubility of each compound in the units given.

11. PbC_2O_4 $Ksp = 2.74 \times 10^{-11}$ (mol/L)
12. $AgBrO_3$ $Ksp = 5.77 \times 10^{-15}$ (g/L)
13. CuS $Ksp = 8.5 \times 10^{-45}$ (mol/L)
14. $CaSO_4$ $Ksp = 2.45 \times 10^{-5}$ (g/