

3 - Molecules & Compounds

Mass Calculations - Difficulty Level 3

1 mole = 6.02 x 10²³ molecules = 22.4 L (at STP)

1. Calculate the mass of 1.28 moles CO₂. (molar mass CO₂ = 44.0 grams)

- (a) 1.28 moles CO₂
 (b) 7 g CO₂



2. What volume will 1.28 moles of CO₂ gas occupy at STP?

- (a) 1.28 moles CO₂
 (b) 7 g CO₂



3. How many molecules are there in a 0.00001 mole sample of H₂O?

- (a) 6020000 moles H₂O
 (b) 7 molecules H₂O



4. What mass of CO₂ gas occupies a volume of 100 L, at STP? (molar mass CO₂ = 44.0 grams)

- (a) 100 L moles CO₂
 (b) 7 g CO₂



5. How many molecules are in a 10.0 gram sample of H₂O? (molar mass H₂O = 18.0 grams)

- (a) 10.0 g H₂O
 (b) 7 molecules H₂O



6. What volume will 1.28 x 10²³ molecules of CO₂ occupy at STP?

- (a) 1.28 x 10²³ molecules CO₂
 (b) 7 L

