

Students should read carefully using the high concentrations for the acids. Don't forget to use safety glasses and wear gloves when using acids.

TABLE 1. PREPARATION OF BUFFER SOLUTIONS

Acid	Base

Name: _____

Date: _____

Section: _____

Class: _____

Multiple-choice questions:

- 1. $\text{pH} = -\log[\text{H}^+]$
- 2. $\text{pH} = \log[\text{H}^+]$
- 3. $\text{pH} = \log[\text{H}^+]$
- 4. $\text{pH} = -\log[\text{H}^+]$

5. The molar concentration of the hydroxide ion in a solution is 0.001 M.

- a. The solution is acidic.
- b. The solution is basic.
- c. The solution is neutral.

6. The pH of a solution is 10.5.