

**Covalent compounds- consist of nonmetal only.****If chemical formula starts with H- it's an acid: find its name or formula in the Acids List.****Exception:** H<sub>2</sub>O and H<sub>2</sub>O<sub>2</sub>.**If chemical formula consist of 2 nonmetals only- use the following naming rules.****Ex:** name the following compound -H<sub>2</sub>O<sub>2</sub>

(a) name the first element- unchanged name from PT - hydrogen

(b) for the second element, take **its root** and add ending **-ide** - **oxide**

(c) add prefixes which will show how many atoms of each element are here in a compound

Answer: **dihydrogen dioxide**Write the names for each of the following covalent compounds.

- |   |  |
|---|--|
| (1) NO - nitrogen monoxide                | (2) HNO <sub>3</sub> - nitric acid                       |
| (3) H <sub>3</sub> PO <sub>4</sub> _____  | (4) N <sub>2</sub> O <sub>3</sub> _____                  |
| (5) SiO <sub>2</sub> _____                | (6) N <sub>2</sub> O <sub>5</sub> _____                  |
| (7) HNO <sub>2</sub> _____                | (8) HBr _____  |
| (9) C <sub>4</sub> H <sub>8</sub> _____   | (10) H <sub>2</sub> SO <sub>3</sub> _____                |
| (11) CCl <sub>4</sub> _____               | (12) HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> _____ |
| (13) HClO <sub>4</sub> _____              | (14) HClO <sub>3</sub> _____                             |
| (15) H <sub>2</sub> SO <sub>4</sub> _____ | (16) H <sub>2</sub> S _____                              |
| (17) HI _____                             | (18) H <sub>2</sub> O <sub>2</sub> _____                 |
| (19) Cl <sub>3</sub> F _____              | (20) N <sub>2</sub> H <sub>4</sub> _____                 |
| (21) Cl <sub>2</sub> O <sub>7</sub> _____ | (22) HCl _____   |
| (23) NH <sub>3</sub> _____                | (24) H <sub>2</sub> SiO <sub>4</sub> _____               |
| (25) Cl <sub>2</sub> O <sub>5</sub> _____ | (26) PCl <sub>3</sub> _____                              |
| (27) HF _____                             | (28) Cl <sub>2</sub> O _____                             |
| (29) CS <sub>2</sub> _____                | (30) H <sub>2</sub> CO <sub>3</sub> _____                |
| (31) SF <sub>4</sub> _____                | (32) SiC _____   |
| (33) P <sub>2</sub> S <sub>3</sub> _____  | (34) F <sub>2</sub> O <sub>5</sub> _____                 |