

Name and Class:

## Mixed Chemical Naming Worksheet

Name these compounds. They may be either ionic or covalent – indicate which are ionic and which are covalent compounds

- 1) LiOH \_\_\_\_\_
- 2) PBr<sub>3</sub> \_\_\_\_\_
- 3) Na<sub>2</sub>SO<sub>4</sub> \_\_\_\_\_
- 4) (NH<sub>4</sub>)<sub>2</sub>S \_\_\_\_\_
- 5) CaCO<sub>3</sub> \_\_\_\_\_
- 6) CF<sub>4</sub> \_\_\_\_\_
- 7) NaNO<sub>3</sub> \_\_\_\_\_
- 8) P<sub>2</sub>S<sub>3</sub> \_\_\_\_\_
- 9) Al(NO<sub>3</sub>)<sub>3</sub> \_\_\_\_\_
- 10) Mg(OH)<sub>2</sub> \_\_\_\_\_

Write the formulas for the following compounds. Remember, they may be either ionic or covalent compounds, so make sure you use the right method!

- 11) potassium oxide \_\_\_\_\_
- 12) phosphorus tribromide \_\_\_\_\_
- 13) calcium hydroxide \_\_\_\_\_
- 14) dinitrogen sulfide \_\_\_\_\_
- 15) carbon monoxide \_\_\_\_\_
- 16) diboron tetrahydride \_\_\_\_\_
- 17) phosphorus pentabromide \_\_\_\_\_
- 18) sulfur dichloride \_\_\_\_\_
- 19) sodium carbonate \_\_\_\_\_
- 20) aluminum acetate \_\_\_\_\_