

Name _____ Date _____ Period _____

Naming Binary Covalent Compounds

COVALENT MOLECULAR COMPOUNDS- COMPOSED OF NONMETALS ONLY.

1. Name of the "more metallic" element is written first. Since both are nonmetals, choose the element closest to the bottom or left of the periodic table.
2. The ending of the second nonmetal is changed to -ide.
3. Prefixes are added to indicate the number of atoms present. Mono is not used on the first element just write the name if there is only one.

1	MONO	2	DI	3	TRI		
4	TETRA	5	PENTA	6	HEXA		
7	HEPTA	8	OCTA	9	NONA	10	DECA

EXAMPLES;

CO is carbon monoxide

CO₂ is carbon dioxide

SF₆ is sulfur hexafluoride

P₂O₃ is diphosphorus trioxide

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|---|--|
| 1. CO _____ | 2. CO ₂ _____ |
| 3. H ₂ O _____ | 4. NH ₃ _____ |
| 5. CH ₄ _____ | 6. NO _____ |
| 7. N ₂ O _____ | 8. N ₂ O ₅ _____ |
| 9. N ₂ O ₃ _____ | 10. PCl ₃ _____ |
| 11. PF ₅ _____ | 12. P ₂ O ₅ _____ |
| 13. SO ₂ _____ | 14. S ₂ O ₇ _____ |
| 15. SiCl ₄ _____ | 16. B ₄ C _____ |
| 17. BN _____ | 18. CS ₂ _____ |
| 19. SeF ₆ _____ | 20. H ₂ O ₂ _____ |
| 21. Cl ₂ O _____ | 22. N ₂ O ₄ _____ |
| 23. NI ₃ _____ | 24. AsCl ₃ _____ |
| 25. CCl ₄ _____ | 26. SeF ₂ _____ |
| 27. SiO ₂ _____ | 28. H ₂ S _____ |
| 29. SF ₄ _____ | 30. SO ₃ _____ |
| 31. XeF ₄ _____ | 32. TeF ₆ _____ |
| 33. BBr ₃ _____ | 34. XeF ₂ _____ |
| 35. Se ₂ Cl ₂ _____ | 36. N ₂ _____ |
| 37. ClF ₃ _____ | 38. BrF ₃ _____ |
| 39. SCl ₂ _____ | 40. S ₂ F ₁₀ _____ |