

Name _____ **answer key** _____ Date _____

Mole Ratio Worksheet

1) Given this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, write the following molar ratios:

a) N_2 / H_2 **1:3**

b) N_2 / NH_3 **1:2**

c) H_2 / NH_3 **3:2**

2) Given the following equation: $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$, write the following molar ratios:

a) $\text{H}_2 / \text{H}_2\text{S}$ **8:8**

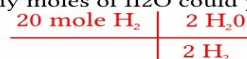
b) H_2 / S_8 **8:1**

c) $\text{H}_2\text{S} / \text{S}_8$ **8:1**

3) Answer the following questions for this equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

a) What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio? **2:2**

b) Suppose you had 20 moles of H_2 on hand and plenty of O_2 . how many moles of H_2O could you make?



c) What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio? **1:2**

d) Suppose you had 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make?



4) Use this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, for the following problems

a) If you used 1 mole of N_2 , how many moles of NH_3 could be produced? **2**



b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required?

c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?



d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?

