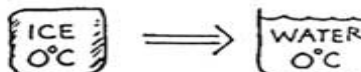


CONCEPTUAL Physics PRACTICE PAGE

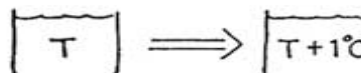
Chapter 17 Change of Phase
Ice, Water, and Steam

All matter can exist in the solid, liquid, or gaseous phases. The solid phase exists at relatively low temperatures, the liquid phase at higher temperatures, and the gaseous phase at still higher temperatures. Water is the most common example, not only because of its abundance but also because the temperatures for all three phases are common. Study "Energy and Changes of Phase" in your textbook and then answer the following:

1. How many calories are needed to change 1 gram of 0°C ice to water?



2. How many calories are needed to change the temperature of 1 gram of water by 1°C?



3. How many calories are needed to melt 1 gram of 0°C ice and turn it to water at a room temperature of 23°C?



4. A 50-gram sample of ice at 0°C is placed in a glass beaker that contains 200 g of water at 20°C.

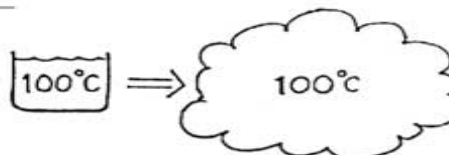


- a. How much heat is needed to melt the ice? _____

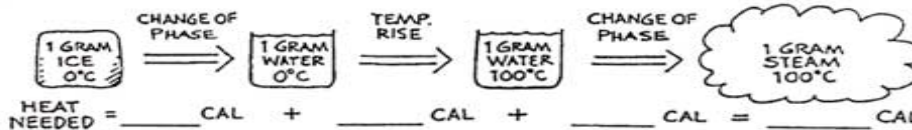
- b. By how much would the temperature of the water change if it gave up this much heat to the ice? _____

- c. What will be the final temperature of the mixture? (Disregard any heat absorbed by the glass or given off by the surrounding air.) _____

5. How many calories are needed to change 1 gram of 100°C boiling water to 100°C steam?



6. Fill in the number of calories at each step below for changing the state of 1 gram of 0°C ice to 100°C steam.



Draw it!