

## Periodic Table Worksheet

Name: \_\_\_\_\_ Date: \_\_\_\_\_

1. Periodic means \_\_\_\_\_  
Examples of periodic properties:
2. What is a group (or family)? \_\_\_\_\_ What is a period? \_\_\_\_\_
3. How can you determine the number of electrons in an element's outer energy level by the group it's in?
4. What is the valence shell?
5. Why do elements that make positive ions appear on the left side of the periodic table while those that make negative ions appear on the right?
6. What is the common name for group 1A? \_\_\_\_\_  
Why do the elements of this group usually not form ions?

7. Complete the following table.

Group	Common Name	Charge on Ions of the Group
1		
2		
13, 14, 15A		
16, 17, 18A		
17, 18, 19A		

8. Predict the charges on ions of the following atoms:

K<sup>+</sup>, \_\_\_\_\_ N<sup>3-</sup>, \_\_\_\_\_ O<sup>2-</sup>, \_\_\_\_\_ Cl<sup>-</sup>, \_\_\_\_\_ In<sup>+</sup>, \_\_\_\_\_ Br<sup>-</sup>, \_\_\_\_\_ F<sup>-</sup>, \_\_\_\_\_

9. a) In group II, which element is the most active? \_\_\_\_\_

b) Metallic activity tends to decrease, direction is one goes down Group II.

10. a) Which element is most active in group 17? \_\_\_\_\_

b) Nonmetal activity tends to increase, direction is one goes down Group 17.

11. Compare and contrast ionization energy and atomic radius.

Ionization Energy	Radius
Definition: _____ _____	_____
Largest values trend or general idea: _____	_____

  

Largest values trend or general idea: _____	_____
_____	_____