

CHEMISTRY**ATOMIC STRUCTURE PRACTICE I**

X = element symbol

A = mass number [# of protons (p) + # neutrons (n)]

Z = atomic number [# of protons]

N = # of neutrons

A - Z = N

A typical isotopic symbol takes this form:



ex. The isotopic symbol for Fluorine would be



Fill in the missing items in the table below.

Name	Symbol	Z	A	#p	#e	#n	Isotopic Symbol
	Na						
		17					
Potassium							

Fill in the missing items in the table below.

Name	Symbol	Z	A	#p	#e	#n	Isotopic Symbol
	P						
Iron							
				53			

Fill in the missing items in the table below.

Name	Symbol	Z	A	#p	#e	#n	Isotopic Symbol
Silver							
		36					
	W						