

Types of Reactions

- _____ 01. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- _____ 02. $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- _____ 03. $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
- _____ 04. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 05. $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$
- _____ 06. $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- _____ 07. $\text{H}_2\text{SO}_4 + \text{AgNO}_3 \rightarrow \text{HNO}_3 + \text{Ag}_2\text{SO}_4$
- _____ 08. $\text{NaCl} + \text{KNO}_3 \rightarrow \text{KCl} + \text{NaNO}_3$
- _____ 09. $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- _____ 10. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 11. $\text{H}_2 + \text{MgO} \rightarrow \text{Mg} + \text{H}_2\text{O}$
- _____ 12. $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
- _____ 13. $\text{H}_2\text{SO}_4 + \text{CaCO}_3 \rightarrow \text{CaSO}_4 + \text{H}_2\text{CO}_3$
- _____ 14. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 15. $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$