

## Types of Reactions

- \_\_\_\_\_ 01      $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
- \_\_\_\_\_ 02      $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- \_\_\_\_\_ 03      $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
- \_\_\_\_\_ 04      $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$
- \_\_\_\_\_ 05      $\text{H}_2 + \text{O} \rightarrow \text{H}_2\text{O}$
- \_\_\_\_\_ 06      $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
- \_\_\_\_\_ 07      $\text{H}_2\text{SO}_4 + \text{AgNO}_3 \rightarrow \text{HNO}_3 + \text{Ag}_2\text{SO}_4$
- \_\_\_\_\_ 08      $\text{MgCl}_2 + \text{KNO}_3 \rightarrow \text{MgNO}_3 + \text{KCl}$
- \_\_\_\_\_ 09      $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
- \_\_\_\_\_ 10      $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- \_\_\_\_\_ 11      $\text{H}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 12      $\text{CaCl}_2 + \text{H}_2 \rightarrow \text{CaH}_2 + \text{HCl}$
- \_\_\_\_\_ 13      $\text{H}_2\text{SO}_4 + \text{CaCO}_3 \rightarrow \text{CaSO}_4 + \text{H}_2\text{CO}_3$
- \_\_\_\_\_ 14      $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- \_\_\_\_\_ 15      $\text{H}_2 + \text{H}_2 \rightarrow \text{H}_2$