

NAME \_\_\_\_\_

HOUR \_\_\_\_\_

APC Ch. 8 Supplement Worksheet: Periodic Trends

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
3. What is the difference between electron affinity and ionization energy?
4. Why does fluorine have a higher ionization energy than iodine?
5. Which element is the smallest: Chlorine, selenium, or bromine?
6. Which atom has the highest ionization energy: phosphorus, sulfur, or selenium?
7. Which is the largest: a potassium atom, a potassium ion, or a rubidium atom?
8. Which is the largest: a chlorine atom, a chlorine ion, or a bromine atom?
9. Arrange the following in order of decreasing radius: Si, Na, Mg
10. Which would have a smaller first ionization energy: oxygen or sulfur? Why?