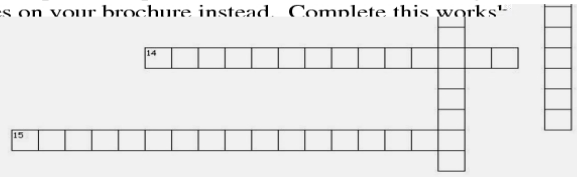
Chemistry
Periodic Table

Name: _____

Period: 1 2 3 4 5 6 7 8

Periodic Table Study Guide

Directions: Please use this packet as practice and review. DO NOT try to answer these questions during presentations, take notes on your brochure instead. Complete this works



Across

2. In cellular respiration, series of anaerobic chemical reactions in the cytoplasm that break down glucose into pyruvic acid; forms a net of two ATP molecules.
10. Chemical process where mitochondria break down food molecules to produce ATP; the three stages are glycolysis, the citric acid cyc the electron transport chain.
12. Molecules that absorb specific wavelength of sunlight.
13. In cellular respiration, series of reactions that break down glucose and produce ATP; energizes electron carriers that pass energize electrons on to the electron transport chain.
14. Process by which autotrophs, such as algae and plants, trap energy from sunlight with chlorophyll and us this energy to convert carbo dioxide and water into simple sugars.
15. Series of proteins embedded in a membrane along which energized electrons are transported; as electrons are passed from molecule molecule, energy is released.

Down

1. Anaerobic process where cells convert pyruvic acid into carbon dioxide and ethyl alcohol; carried out by many bacteria and fungi such yeasts.
3. Energy-storing molecule in cells composed of an adensosine molecule, a ribose sugar and three phoshate groups; energy is stored in th molecules's chemical bonds and can be used quickly and easily by cells.
4. Phase of photosynthesis where light energy is converted to chemical energy in the form of ATP; results in the splitting of water and release of oxygen.
5. Electron carrier molecule; when carrying excited electrons, it becomes NADPH.
6. Reaction taking place in the thylakoid membranes of a chloroplast during the light-dependent reactions where t

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ay help you.

5- Pent
6- Hex
7- Sept
8- Oct
9- Ein

0- Nil
1- Un
2- Bi
3- Tri
4- Quad

Unnilquadium? _____

es that the physical and chemical properties of elements are periodic
b) atomic numbers, (c) radii, (d) structures.

follow the periodic law except: (a) valance electrons, (b) electron
d physical properties, (d) atomic number, (e) symbols and names.

masses of Te and I decrease rather than increase, while their atomic
henon happens to other neighboring elements in six other places on the
e sets of elements, more for bonus points.

7. What is the atomic number of U

8. _____ The periodic law sta functions of their: (a) masses, (

9. _____ All of the following configurations, (c) chemical an

10. In the periodic table, the atomi numbers increase. This phenor periodic table. Name two of th

Alkali/Alkaline Earth Metals

and the alkaline earth metals **green** on the periodic at the back of the
m of the page that identifies the meaning of each color. Continue

1. Color the alkali metals **orange** packet. Draw a key at the bott