

Name: \_\_\_\_\_

Date: \_\_\_\_\_

**Module 10**

Students must answer these questions in a separate sheet. Additional student answers may be found on your personal sheet.

Example:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  The mass of 2 moles of hydrogen reacting with 1 mole of oxygen is

$2 \times 2 \text{ g} + 32 \text{ g} = 36 \text{ g}$

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$4 \text{ g} + 32 \text{ g} = 36 \text{ g}$

3.6 g of water

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→ The total mass of reactants equals the total mass of products.

1.  $\text{CaO}$  \_\_\_\_\_

11.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

2.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

12.  $\text{CaCO}_3$  \_\_\_\_\_

3.  $\text{H}_2\text{O}$  \_\_\_\_\_

13.  $\text{CaSO}_4$  \_\_\_\_\_

4.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

14.  $\text{CaCO}_3$  \_\_\_\_\_

5.  $\text{CaCO}_3$  \_\_\_\_\_

15.  $\text{CaCO}_3$  \_\_\_\_\_

6.  $\text{Na}_2\text{CO}_3$  \_\_\_\_\_

16.  $\text{CaO}$  \_\_\_\_\_

7.  $\text{H}_2\text{O}$  \_\_\_\_\_

17.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

8.  $\text{CaO}$  \_\_\_\_\_

18.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

9.  $\text{CaO}$  \_\_\_\_\_

19.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

10.  $\text{CaCO}_3$  \_\_\_\_\_

20.  $\text{H}_2\text{O}$  \_\_\_\_\_

21.  $\text{H}_2\text{O}$  \_\_\_\_\_

21.  $\text{AgF}$  \_\_\_\_\_

22.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

22.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

23.  $\text{CaF}_2$  \_\_\_\_\_

23.  $\text{H}_2\text{O}$  \_\_\_\_\_

24.  $\text{H}_2\text{O}$  \_\_\_\_\_

24.  $\text{CaCO}_3$  \_\_\_\_\_

25.  $\text{H}_2\text{O}$  \_\_\_\_\_

25.  $\text{H}_2\text{O}$  \_\_\_\_\_

26.  $\text{CaCO}_3$  \_\_\_\_\_

26.  $\text{H}_2\text{O}$  \_\_\_\_\_

27.  $\text{CaO}$  \_\_\_\_\_

27.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

28.  $\text{CaO}$  \_\_\_\_\_

28.  $\text{H}_2\text{O}$  \_\_\_\_\_

29.  $\text{CaO}$  \_\_\_\_\_

29.  $\text{CaCO}_3$  \_\_\_\_\_

30.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_

30.  $\text{CaCO}_3$  \_\_\_\_\_